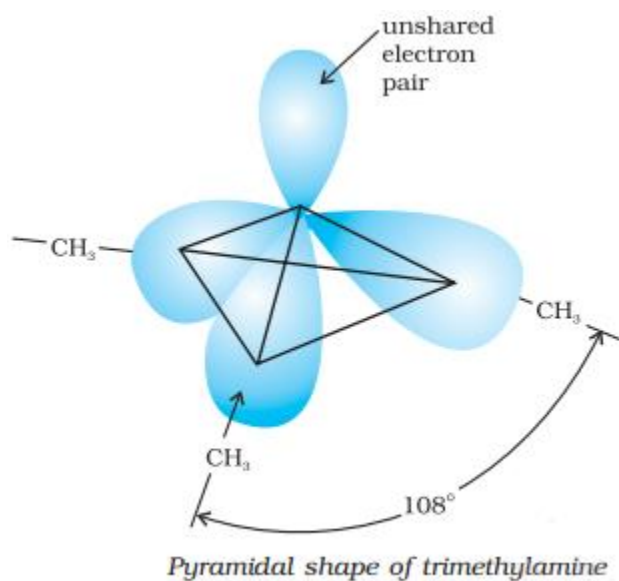


## Amines

**Structure of amines:** Trigonal Pyramidal



## Basicity of Amines

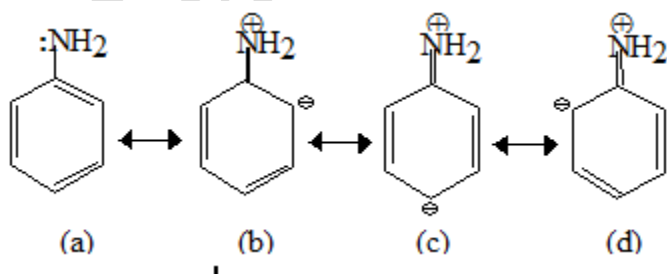
### a) Aliphatic Amine:

Basic Strength:  $\text{NH}_3 < \text{RNH}_2 < \text{R}_2\text{NH} < \text{R}_3\text{N}$ ,

### b) Aromatic Amine:

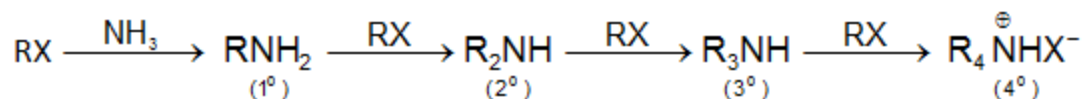
Basic Strength:  $\text{NH}_3 > \text{Ar-NH}_2 > \text{Ar}_2\text{-NH}_2$

Benzene ring decreases the electron density over N atom due to resonance effect.

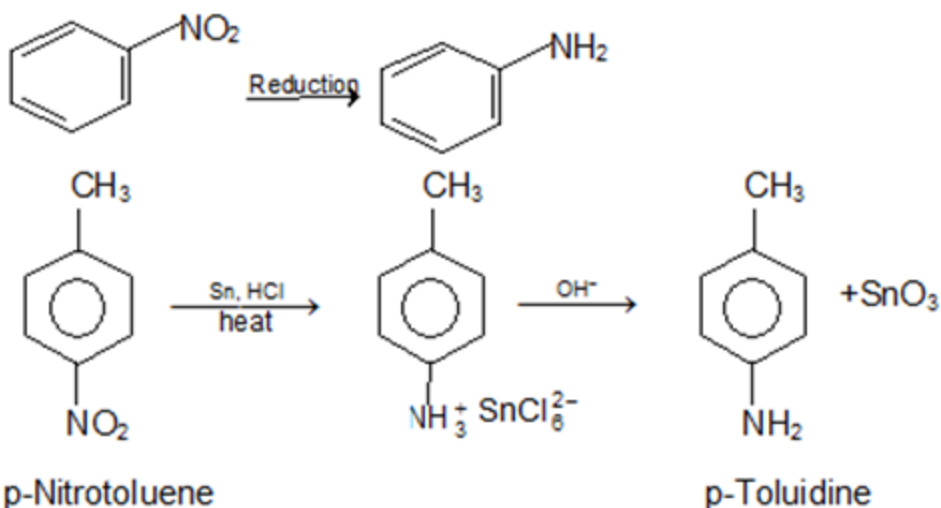


## Preparation of Amines

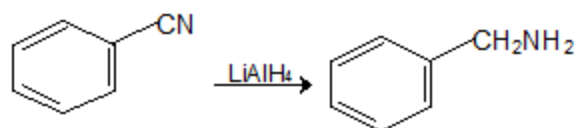
**a) From Alkyl halides:**



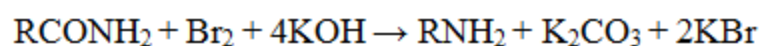
**b) From Nitro Compounds:**



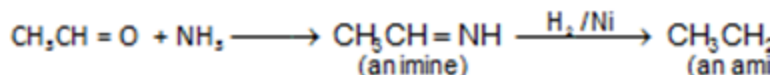
**c) From Nitriles :**



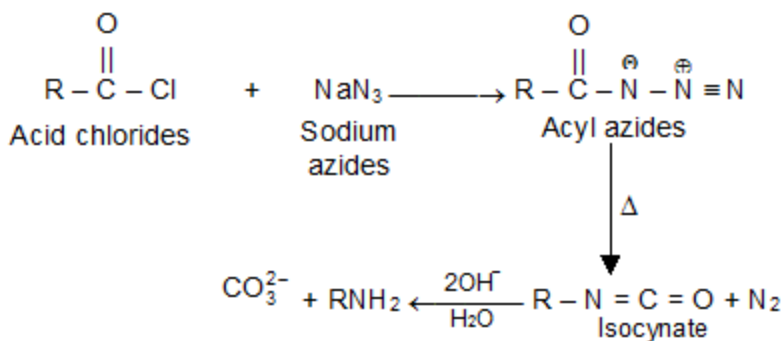
**d) Hofmann Bromamide or Hofmann degradation:**



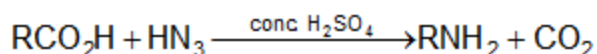
**e) From carbonyl compounds (Reductive Amination)**



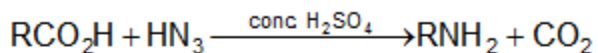
**f) Curtius reaction:**



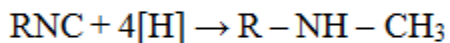
**g) Schmidt reaction :**



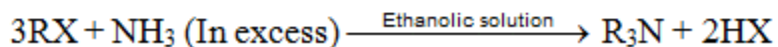
**g) Schmidt reaction :**



**h) Reduction of Alkyl isocyanide:**

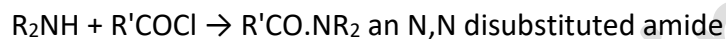


**i) Preparation of tertiary amine:**



**Chemical Reactions**

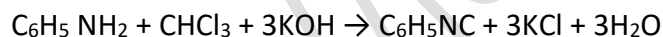
**a) Acylation:**



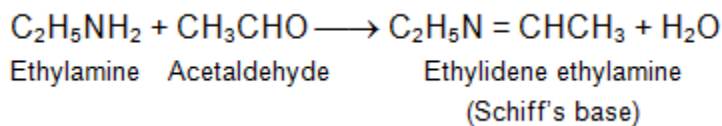
**b) Benzoylation (Schotten Baumann Reaction)**

Primary amine reacts with benzoyl chloride to give the acylated product

**c) Carbylamine Reaction (Given Only by Primary Amines):**

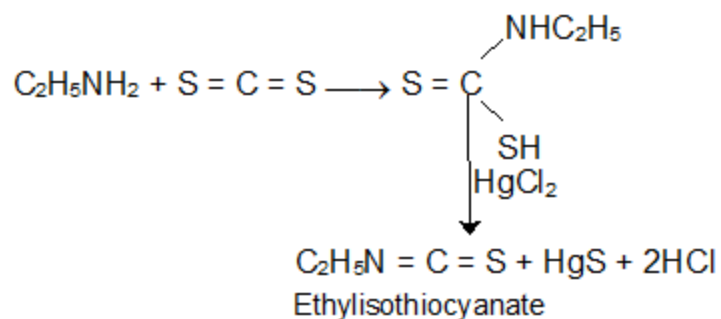


**d) Action with Aldehyde and Ketone:**



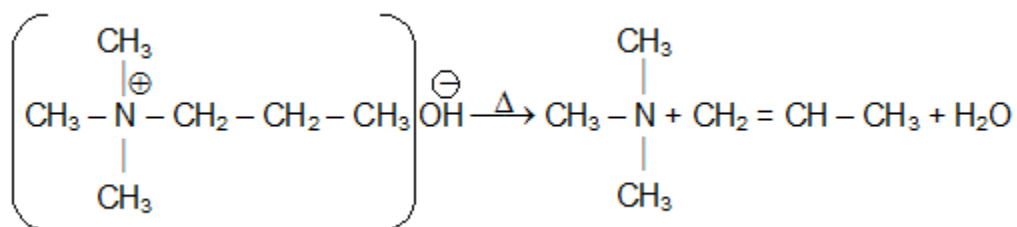
**e) Hofmann Mustard Oil Reaction:**

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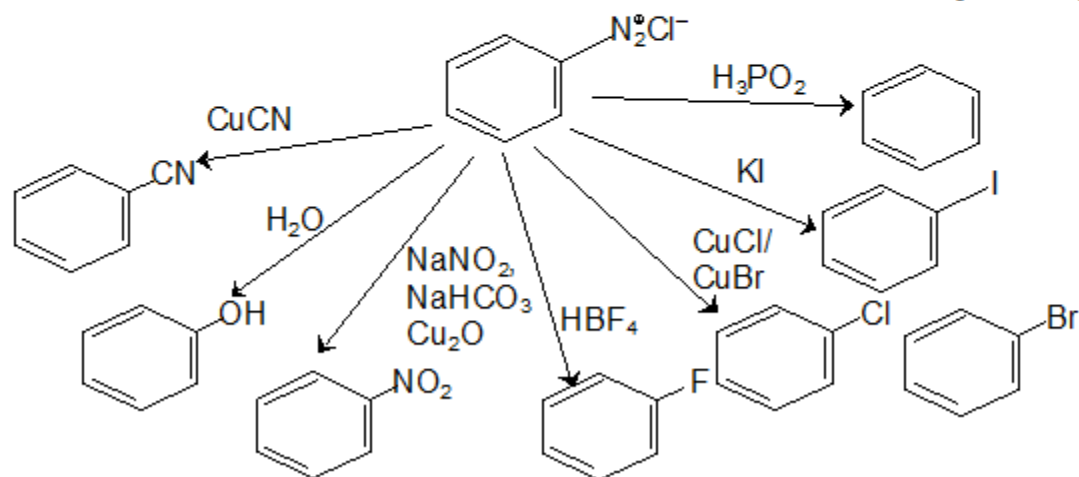
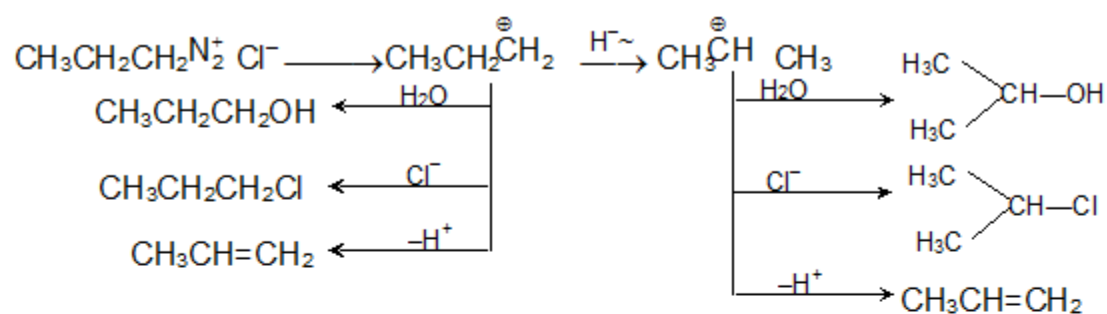
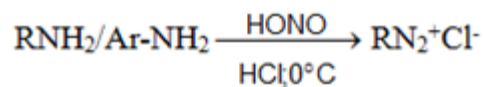


f) **Reaction with Carbonyl Chloride:**  $\text{C}_2\text{H}_5 - \text{NH}_2 + \text{COCl}_2 \rightarrow \text{C}_2\text{H}_5\text{NCO} + 2\text{HCl}$

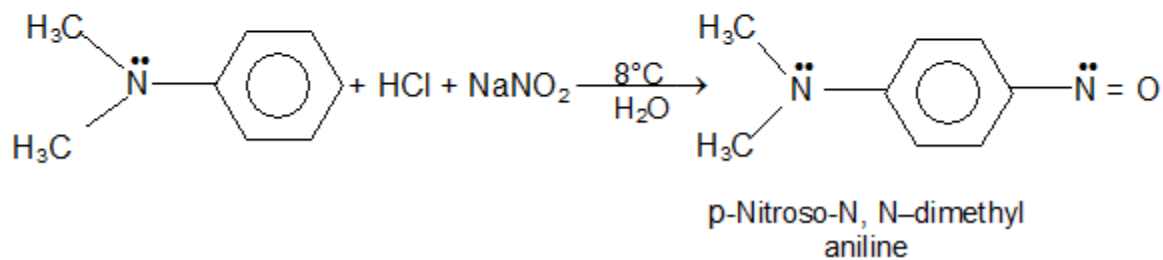
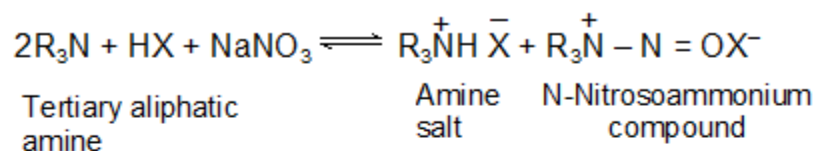
g) **Hofmann Elimination:** When a quaternary ammonium hydroxide is heated strongly ( $125^\circ$  or higher) it decomposes to yield water, a tertiary amine and an alkene



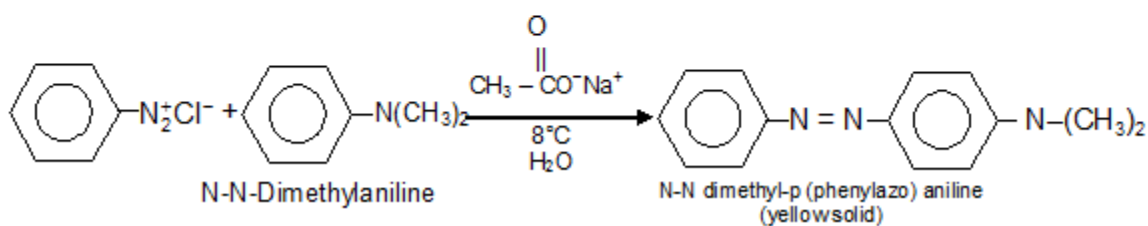
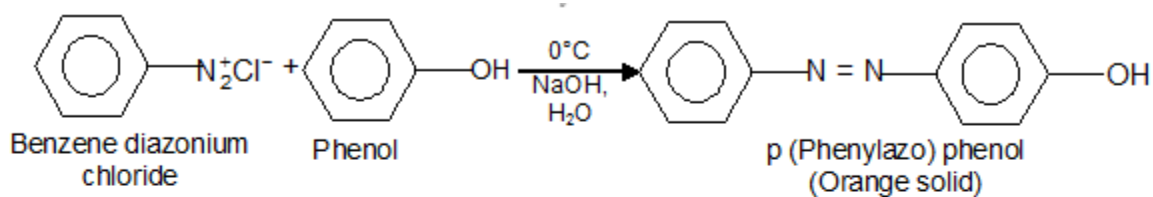
h) **The diazonium salts of amines:**



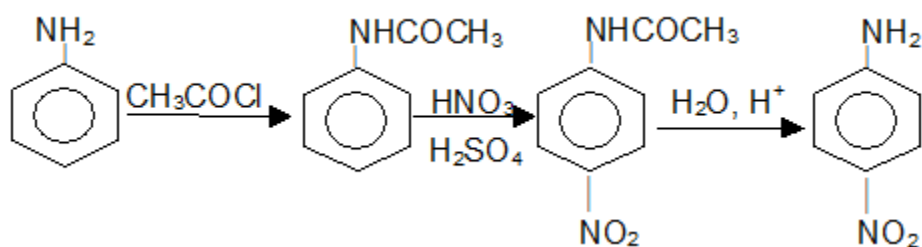
**i) Reaction of Tertiary amines with Nitrous acid:** When a tertiary aliphatic amine is mixed with nitrous acid, an equilibrium is established among the tertiary amine, its salt, and an N-Nitrosoammonium compound.



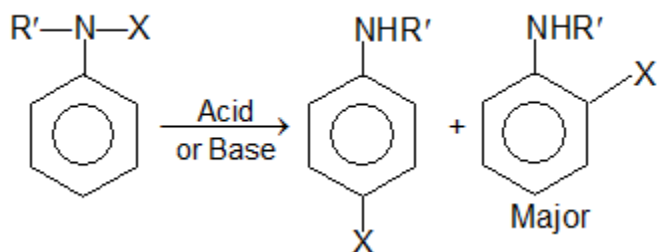
**j) Coupling Reactions of Arene Diazonium Salts:**



**k) Ring Substitution in Aromatic Amines:**



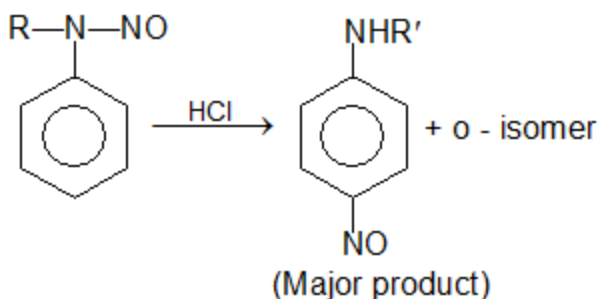
**l) Aniline -X rearrangement:**



Such compounds are not much stable so the group X migrates mainly at p-position.

**1. Fisher-Hepp rearrangement**

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2. Phenylhydroxylamine - p-aminophenol rearrangement.

Separation of a Mixture of Amines:

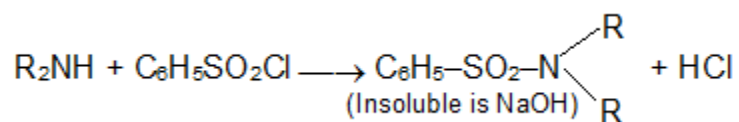
a) Hinsberg's Method

**Primary amine:**  $\text{RNH}_2 + \text{C}_6\text{H}_5\text{SO}_2\text{Cl} \xrightarrow{\text{H}^+} \text{C}_6\text{H}_5\text{-SO}_2\text{-NH-R} + \text{HCl}$

$\text{C}_6\text{H}_5\text{-SO}_2\text{-NH-R}$  : N-alkyl benzene sulfonamides

Dissolves in NaOH due to acidic H-attached to Nitrogen)

**Secondary amine**

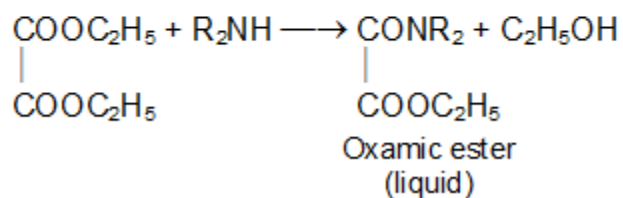
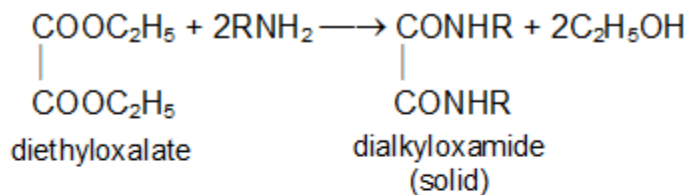


**Tertiary amine** : Tertiary amines do not react with Hinsberg's reagent.

a) Hofmann's Method:

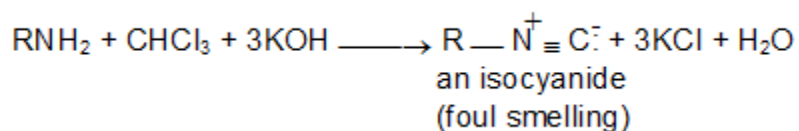
The mixture of amines is treated with diethyloxalate, which forms a solid oxamide with primary amine, a liquid oxime ester with secondary amine. The tertiary amine does not react.

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### Test for Amines

**Primary amine** is treated with a strong base in presence in chloroform, an isocyanide is formed and this isocyanide thus formed has a very foul smell.



**Secondary amine** is converted into nitrosoamine by treating the amine with nitrous acid. The resultant solutions warmed with phenol and concentrated  $\text{H}_2\text{SO}_4$ , a brown or red colour is formed at first soon it changes to blue and then to green. The colour changes to red on dilution and further changes to greenish blue on treating with alkali.

**Tertiary arylamines** react with nitrous acid to form o-nitroso aromatic compound